

Chapter 6 Reaction Practice

For each of the following reactions, write a balanced equation for the reaction. Coefficients should be in terms of lowest whole numbers.

1. Magnesium ribbon is burned in oxygen.
2. A bar of strontium metal is immersed in a 1.0 *M* copper(II) nitrate solution.
3. A sample of nickel(II) sulfite hexahydrate is heated strongly.
4. Solutions of sodium chromate and lead(II) nitrate are mixed.
5. Liquid bromine is carefully added to a solution of potassium iodide.
6. A solution of sodium iodide is added to a solution of lead(II) acetate.
7. Pure solid phosphorus (P_4) is burned in air.
8. Solid sodium carbonate is strongly heated.
9. A piece of copper wire is placed in a solution of silver nitrate.
10. Solutions of strontium nitrate and sodium sulfate are mixed.

11. A small piece of calcium metal is added to hot distilled water.
12. Butanol(C_4H_9OH) is burned in air.
13. A solution of copper(II) chloride is added to a solution of sodium sulfide.
14. Solid potassium chlorate is heated strongly.
15. Powdered strontium oxide is added to distilled water.
16. Solutions of cobalt(II) nitrate and sodium hydroxide are mixed.
17. Ethene(C_2H_4) gas is burned in air.
18. A strip of zinc is added to a solution of hydrochloric acid.
19. Solid nickel(II) sulfide is strongly heated in air.
20. Solid pieces of potassium and iodine are heated strongly
21. A piece of nickel metal is immersed in a solution of copper(II) sulfate.
22. A solution of potassium phosphate is mixed with a solution of calcium acetate.