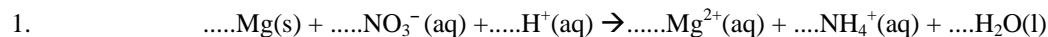


Name _____

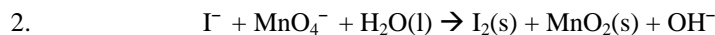
AP Chemistry

___/___/___

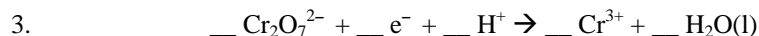
Redox Balancing



When the skeleton equation above is balanced and all coefficients reduced to their lowest whole-number terms. what is the coefficient for H^+ ?



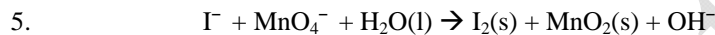
When the skeleton equation above is balanced and all coefficients reduced to their lowest whole-number terms. what is the coefficient for I^- ?



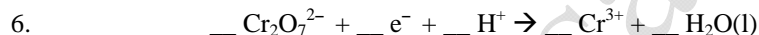
When the equation for the half reaction above is balanced with the lowest whole-number coefficients, the coefficient for H_2O is:



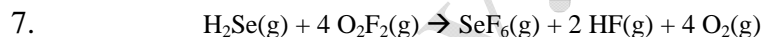
When the skeleton equation above is balanced and all coefficients reduced to their lowest whole-number terms. what is the coefficient for H^+ ?



When the skeleton equation above is balanced and all coefficients reduced to their lowest whole-number terms. what is the coefficient for MnO_4^- ?



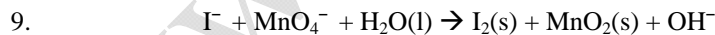
When the equation for the half reaction above is balanced with the lowest whole-number coefficients, the coefficient for $\text{Cr}_2\text{O}_7^{2-}$ is:



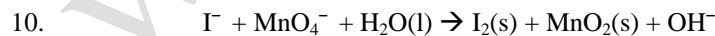
Which of the above is the reducing agent?



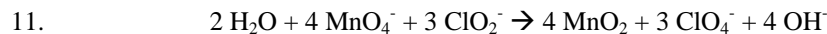
Which of the above is the oxidizing agent?



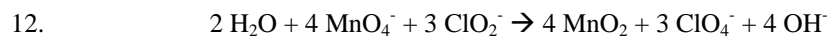
Which of the above is the oxidizing agent?



Which of the above is the reducing agent?



Which of the above is the reducing agent?



Which of the above is the oxidizing agent?